**Lewis electron-dot structures**

* Lewis electron-dot structures are \_\_\_\_\_\_\_\_\_\_\_\_\_ or \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of atoms and their valence (or outer shell) electrons.
  + Electron dot structures show only the electrons in the \_\_\_\_\_\_\_\_\_\_ or outer energy level.
* First the atomic symbol is written.
* Then the dots representing the valence electrons are put on each side of the \_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_.
  + Each ­\_\_\_\_\_\_\_ must have at least one dot before dots can be \_\_\_\_\_\_\_\_\_\_\_\_ up.
* Lewis electron-dot structures are used to help determine how \_\_\_\_\_\_\_\_\_\_\_\_ will occur between \_\_\_\_\_\_\_\_\_ and the possible shapes that the molecules will form.
* Sulfur
  + Sulfur has \_\_ valence electrons.

*Practice*

Draw the Lewis electron-dot structure for the following elements.

* Hydrogen
* Phosphorus
* Bromine
* Nitrogen