Unit 1 Review

1. List the steps of the scientific method.
2. Draw and label the chemical hazard label.
3. The formula for density is: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.
4. Define accuracy vs precision.
5. Convert the following to scientific notation:
6. 3892400 c. 0.04529
7. 0.00000023740 d. 102940
8. How many significant figures are in the following numbers:
9. 0.0101 e. 1.00
10. 5200 f. 8.0800
11. 307000 g. 0.0000006
12. 56920200 h. 11.40000
13. Perform the following operations (answer in significant figures):
14. 4.98 x 2.7 = c. 6.43 + 2.984 =
15. 10.98 / 4.37 = d. 12.01 – 10 =
16. A scientist experimentally determined the density of lead to be 8.45 g/cm3. The actual density of lead is 11.342 g/cm3. Calculate the percent error.
17. What is the density of an object having a mass of 8.0 g and a volume of 53.0 ml?
18. What is the volume of an object with a density of 10.5 g/ml and a mass of 25.7 g?