

# ELECTRON CONFIGURATION (LEVEL ONE)

Name \_\_\_\_\_

Electrons are distributed in the electron cloud into principal energy levels (1, 2, 3, ...), sublevels (s, p, d, f), orbitals (s has 1, p has 3, d has 5, f has 7) and spin (two electrons allowed per orbital).

**Example:** Draw the electron configuration of sodium (atomic #11).



Draw the electron configurations of the following atoms.

1. Cl

2. N

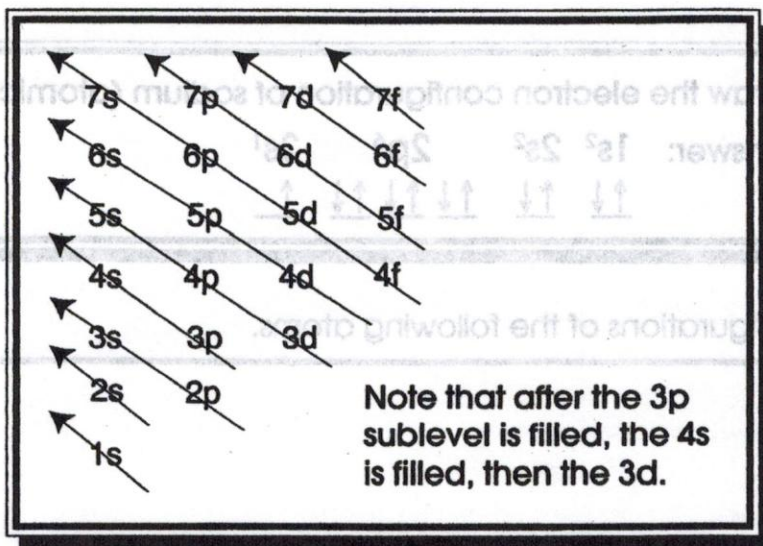
3. Al

4. O

# ELECTRON CONFIGURATION (LEVEL TWO)

Name \_\_\_\_\_

At atomic number greater than 18, the sublevels begin to fill out of order. A good approximation of the order of filling can be determined using the diagonal rule.



Draw the electron configurations of the following atoms.

1. K
2. V
3. Co
4. Zr